

Quantum Thread Theory

by James Cranwell

<http://www.mccelt.com/>

**Everything in the Universe is made from one type of particle.
All workings of the Universe are result from said particle.**

ELECTRONS ARE FIXED POSITION

~~~~~ Bond Angles Are Exact ~~~~~

Electrons must be fixed position -- anything else is impossible.
Notice they way a lot of molecules tend to form tetrahedrally?
A tetrahedron is the opposite diagonal corners of a cube.
And the reason being: electrons form a thread-mesh-type cage around the nucleus. Certain sizes are of course the 8 corners of a cube -- that's what the octet rule is and how it happens.

A lot of atoms and molecules have a bond angle of exactly 109.4712° (degrees).
That's tetrahedral. -- corners of the cube.
That is NOT probability.
It is NOT a cloud.
It is NOT a blur.
It is NOT uncertainty.
It is NOT counterintuitive.

IT IS EXACT.

Electrons DO NOT have any probability or uncertainty involved

Electrons actually are something but everyone mistakenly thinks the vibration travelling around the thread is the electron -- that's what has caused all the confusion.

The electron is conveying vibrations but the material it is made from (quantum threads) are NOT moving..

EXAMPLE: Think guitar string -- the string itself would be the electron but everyone thinks the vibration or note is the electron.

That is why there is all kinds of probability and uncertainty -- the vibration is traveling around a spherical thread mesh cage -- where exactly is the vibration? No way to know for sure.

Got that? The electron has exact position. The vibration position is of course unknown.

SPECIAL NOTE: Heisenberg would be o.k. with this. A vibration of course has uncertain position

~~~~~ The Octet Rule ~~~~~

The reason why there are bonds, bond order and bond angles is because electrons form a thread mesh-type cage around the nucleus -- electrons are fixed positions.

The "Octet Rule" is a perfect 8 corner cubic structure of electrons.

It not only explains everything it makes the explanations an easy feet.

EXAMPLE: They don't know how methane CH_4 can form?

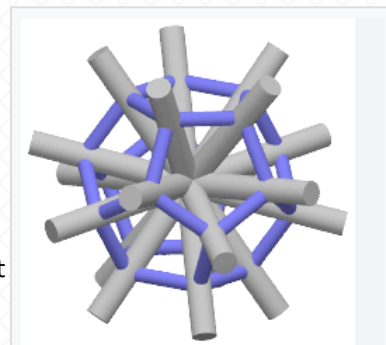
It's an easy answer...

Methane (one carbon atom and 4 hydrogen atoms) CH_4 is the greatest example of fixed position electrons.

There would be a cubic carbon atom and 4 corners would each have an hydrogen atom electron fixed into place.

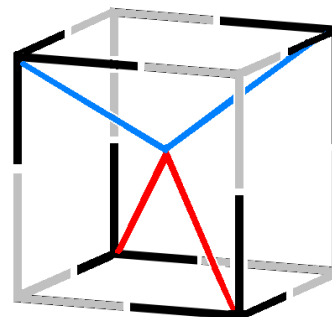
If the picture is thought of as a carbon atom -- four electrons are fixed to the

nucleus. The other four (grey) empty spots can accept electrons from other atoms -- in this case hydrogen.



The particle itself would be just the grey threads (or strings) in the picture (no color and a lot thinner of course).
It would fit perfectly inside of a **dodecahedron**.
Actual thread (or string) length is about one Ångström and it is fine enough were 10 threads (20 radii) could curl-up into the size of a neutron.

Octet Rule and Bond Angles



109.4712° (degrees)

~~~~~  
**The Thread Mesh-Type Cage**  
 ~~~~~

Larger sized elements are unstable because electrons form a thread-mesh-type cage of concentric spheres around the nucleus.
 The basic "electron" is made from 10 threads joined at their centers (that forms 20 radii emanating from a common center).
 The threads are approx. one angstrom in length.
 One thread is attached to the proton, 18 form the electron "disc" and the other thread connects to outer electrons. That's why the periodic table has 18 columns -- when electrons join each other they have 18 threads to play with. Got that?

Here are two electrons...

~~~~~\*~~~~~      ~~~~~\*~~~~~

When the atom is small (inner shells) threads can completely twist and connect together.

~~~~~\*≈≈≈≈≈\*~~~~~

But as it gets larger (outer shells) the electrons might be further apart and therefore less thread connection twist

~~~~~\*~~~~~≈≈≈~~~~~\*~~~~~

until the point of instability...

~~~~~\*~~~~~≈~~~~~\*~~~~~

~~~~~  
**Every Corner Is An Electron**  
 ~~~~~

NOTE: An electron thread goes from the nucleus to the electron disc. There the disc forms one corner of the cube. An electron disc has 18 threads, but they do not reach diagonally across the cube, so the 18 threads bunch together into three groups of 6 threads. All corners actually have 18 threads divided equally and attaching to other corners.
 A cube is an automatic shape that forms in the "shell" because the electron threads attach to other electrons based on the amount of electrons and the distance from nucleus -- they shoot for a tight around-the-sphere pack.
 EXAMPLE: If you opened-up 8 umbrellas; they would form an umbrella ball -- something like the spherical cube pack. But if you increased the length of the handles you would need a lot more umbrellas to make and fill-in an umbrella sphere -- that would be like an outer shell of an atom.
 Any other configuration you can find or think of also proves the electrons are fixed position. It depends on the number of electrons, what shell they are in, if there are double bonds, etc.
 QUESTION: If you have an atom or molecule and everything is bonding trigonally at exactly 120° (degrees) ...would you say the electrons are orbiting randomly or do they have a definite fixed position?

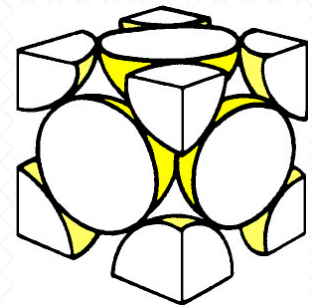
Face-centered cubic, body-centered cubic, all crystalline structures. Nothing is random -- everything is exact.

So, Yes -- electrons are fixed position. And now you know how they work.
 Here is a regular thread tension formula...

Tension = velocity squared x mass / Length.

If we plug c in and rearrange we get the one-inch formula... $TL = mc^2$

<http://www.mccelt.com/the-one-inch-equation-to-explain-all-physical-laws.php>



References

[3] Quantum Thread Theory & Why the Speed of Light is "C"
<http://vixra.org/abs/1612.0363>
 Authors: Seamus McCelt
 Category: Quantum Gravity and String Theory

Electrons are fixed position

Octet Rule and Bond Angles

109.4712° (degrees).

That's tetrahedral.

Corners of the cube.

That is NOT probability.

It is NOT a cloud.

It is NOT a blur.

It is NOT uncertainty.

NOT counterintuitive.

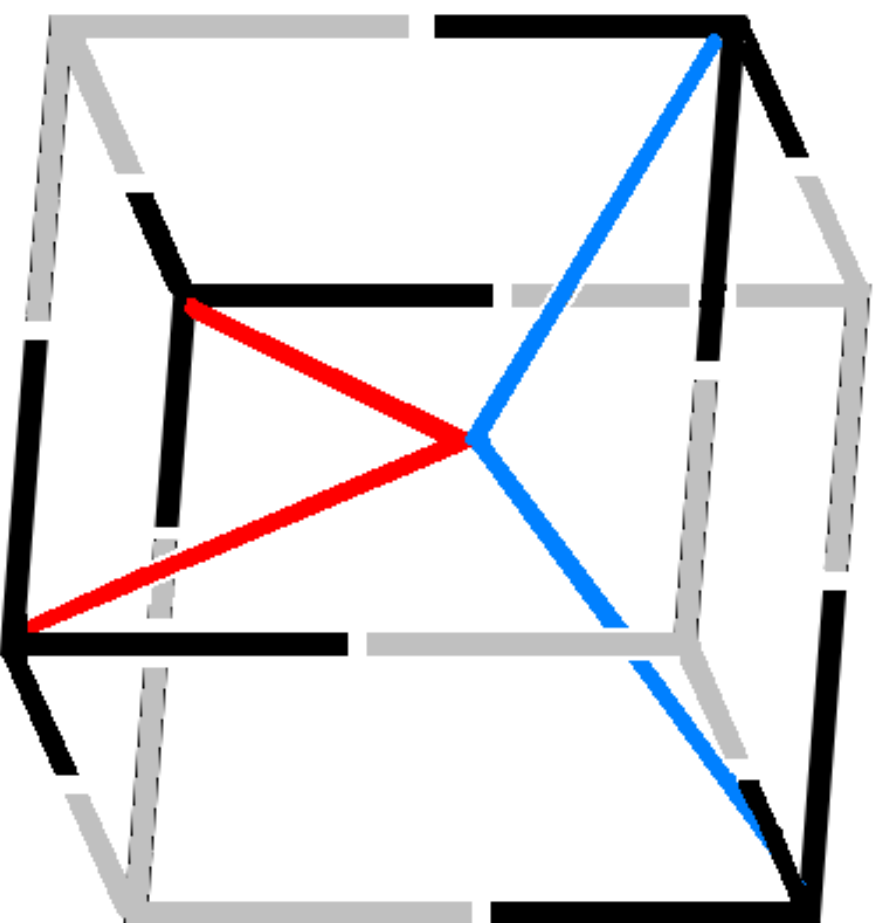
IT IS EXACT.

Electrons DO NOT

have any probability

or uncertainty involved.

mccelt.com



109.4712° (degrees)

$E_L = mc^2$